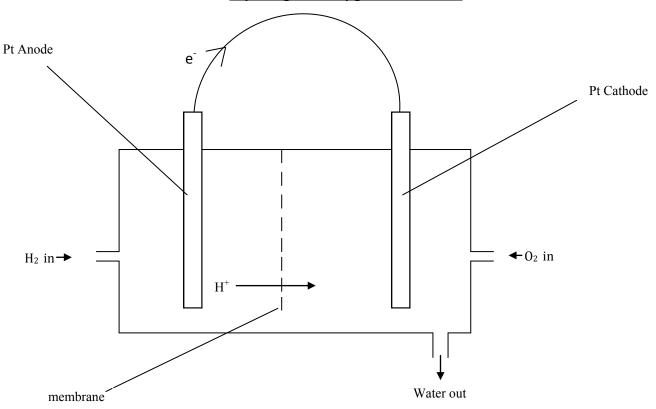
Hydrogen-Oxygen Fuel Cell



$$2H_2 \xrightarrow{\underline{\text{Pt Anode}}} 4H^+ + 4e^-$$

To Remember

Look at me and think

Old Age X n i o d at d a e t i o n

Oxidation

S

Loss

Reduction

Is

Gain

Adding Anode and Cathode reactions:

$$2H_2 \xrightarrow{\bullet} 4H^{\bullet} + 4e^{\bullet}$$

$$O_2 + 4H^{\bullet} + 4e^{\bullet} \xrightarrow{\bullet} 2H_2O$$

Overall: $2H_2 + O_2 \longrightarrow 2H_2O$